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oxygen are required to react completely with 4.0 liters of hydrogen to form water? $2\text{H}_2(\text{g}) + \text{O}_2 \rightarrow 2\text{H}_2\text{O}(\text{g})$ 2.0 L.

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sylvviewiley. Terms in this set (14)
stoichiometry. the study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction. mole ratio.

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15.2 CHAPTER 11: STOICHIOMETRY.

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MOLE TO MOLE RATIO. When nitrogen and hydrogen gas are heated under the correct conditions, ammonia gas (NH_3) is formed. a. RXN: $1. \text{N}_2 + 3. \text{H}_2 \rightarrow 2. \text{NH}_3$. b. How many moles of nitrogen react with three moles of hydrogen? 1
mol N_2 3 mol H_2 1 mol N_2 . 3 mol H_2 . c. How many moles of nitrogen react with six moles ...

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Section 11.1 • Defining Stoichiometry
371 Mole ratios You have read that the coefficients in a chemical equation indicate the relationships between moles of reactants and products.

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Stoichiometry Stoichiometry CHAPTER 11
SOLUTIONS MANUAL Section 11.1

Defining Stoichiometry pages 368–372

Practice Problems pages 371–372 1.

Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that

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the law of conservation of mass is

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Define stoichiometry. What does
stoichiometry do for you? How is the
molar ratio determined? What do

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coefficients represent? In a mass-mass problem, mass is first converted to what? How is number of moles found from mass and molar mass? What are the general steps (conversions) for a mass-mass problem?

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units and chemical formulas. ...

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13. 11.2 g of metal carbonate, containing an unknown metal, M, were heated to give the metal oxide and 4.4 g CO₂. $MCO_3(s) + \text{heat} \rightarrow MO + CO_2(g)$... Practice Test Ch3 Stoichiometry (page 3 of 3) 1. d It might be easiest to balance the

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equation with mostly whole numbers: $2\text{NH}_3 + \frac{7}{2}\text{O}_2 \rightarrow 2\text{NO}_2 + 3\text{H}_2\text{O}$. The question asks

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Chapter 11: Stoichiometry During a chemical reaction, reactants are consumed as new products are... explore. take Self-Check Quizzes. projects, and activities. find the Try at. Stoichiometry is the tool for answering these questions. and solutions.

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